

SYNTHESIS OF HETEROCYCLIC THIOAMIDES AND COPPER(II)  
COORDINATION COMPOUNDS BASED ON THEMAnatoliy Ranskiy<sup>1,\*</sup>, Natalia Didenko<sup>1</sup>, Olga Gordienko<sup>1</sup>

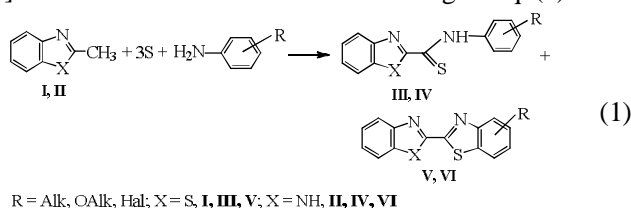
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**Abstract.** Arylamides of hetaryl-2-thiocarboxylic acid were synthesized *via* Willgerodt–Kindler modified reaction. The final products were studied by means of chemical analysis, IR- and <sup>1</sup>H NMR-spectroscopy. The coordination compounds of the general formula [Cu(HL<sup>4</sup>)Cl<sub>2</sub>]<sub>2</sub> · xCH<sub>3</sub>OH (x = 0, 2) were obtained *via* traditional and direct synthesis based on benzimidazol-2-N-(4-ethoxyphenyl) carbothioamide (HL<sup>4</sup>). Using X-ray analysis the molecular and crystal structures of [Cu(HL<sup>4</sup>)Cl<sub>2</sub>]<sub>2</sub> · 2CH<sub>3</sub>OH complex were determined.

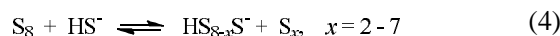
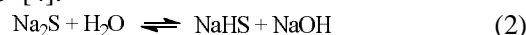
**Keywords:** heterocyclic thioamide, copper(II) binuclear complex, X-ray analysis.

## 1. Introduction

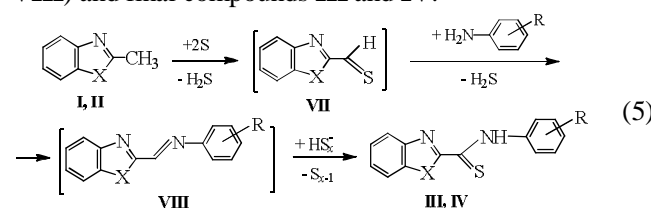
The modern development of coordination chemistry of 3d-metals and N,S-containing organic ligands is mainly connected with the synthesis of compounds with new biological [1], catalytic [2] and physico-mechanical [3] properties. The synthesis of heterocyclic thioamides with benzothiazole and benzimidazole fragments is complicated and connected with investigations of Willgerodt–Kindler modified reaction [4]. It is the reason for searching new methods of thioamides synthesis, their effective using in the coordination chemistry of 3d-metals [5] and obtaining compounds with a set of new chemical and physical properties. Previously the heterocyclic thioamides were obtained *via* thermal [6] or catalytic [3, 4] activation of elemental sulfur according to Eq. (1):



The by-products of oxidative cyclization **V** and **VI** are formed additionally to the main arylamides of hetaryl-2-thiocarboxylic acid **III** and **IV**. In case of catalytic reaction, the sulfur-containing nucleophilic catalyst (Na<sub>2</sub>S·9H<sub>2</sub>O) is used. The catalyst decomposes the initial cyclooctasulfan S<sub>8</sub> to the reactive nucleophilic ions HS<sup>-</sup> and HS<sub>8-x</sub>S<sup>-</sup> [4]:



The oxidation of active methyl group of the initial compounds **I** and **II** by sulfur (intermediate **VII**) is accompanied by formation of Schiff bases (intermediate **VIII**) and final compounds **III** and **IV**.



It should be noted that using the catalyst the yield of heterocyclic thioamides is increased by 2.0–2.5 times [3, 4] compared with the reaction thermal activation (Eq. (1)). However, in this work we propose to improve the synthesis of heterocyclic thioamides by adding dimethylsulfoxide (DMSO) that increases their yield and decreases the reaction temperature. The synthesized thioamides were analyzed as bidentate ligands while complex formation with some 3d-metals.

It should be also noted that nowadays the direct synthesis of coordination compounds of transition 3d-metals and thioamide ligands using HL–ROH–HCl–O<sub>2</sub> proton-donor systems is practically unstudied [7]. Meanwhile, such coordination compounds are widely used in engineering as additives to hydrocarbon materials [8] and model objects of copper(II) [8] and cobalt(II, III) [9] biologically active complexes. In addition, just theoretical studies of the mechanism of the redox reactions occurred on the M<sup>0</sup>-Ox-HL-Solv interface [10], are of particular interest.

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So, the aim of this work is to investigate copper(0, II) complex formation with heterocyclic thioamides.

## 2. Experimental

### 2.1. Materials and Methods

To synthesize the heterocyclic thioamides **IIIa–IIIi** (Table 1) and **IVa–IVg** (Table 2) we used 2-methylbenzothiazole, 2-methylbenzimidazole and aromatic amines purchased by Aldrich and Merck. The compounds were of CP grade with the main component content of 99 %.

For the synthesis of  $[\text{Cu}(\text{HL}^4)\text{Cl}_2]_2 \cdot 2\text{CH}_3\text{OH}$  complex we used the copper powder with mass part of metal  $\geq 99.5\%$  and particles size of  $8.0 \pm 1.1 \mu\text{m}$  (75 vol %), determined by Saishin SKC-2000S microsedimentometer (Japan). Inorganic salts  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$  and  $\text{Na}_2\text{S} \cdot 9\text{H}_2\text{O}$  were of CP grade. Organic solvents (*i*- $\text{C}_3\text{H}_7\text{OH}$ ,  $\text{CH}_3\text{OH}$  and DMSO) were purchased by Merck and Aldrich and used as received without additional purification. Copper(II) content in the synthesized complex  $[\text{Cu}(\text{HL}^4)\text{Cl}_2]_2 \cdot x\text{CH}_3\text{OH}$  ( $x = 0, 2$ ) was determined by atomic-absorption spectroscopy using S-115 PKRS spectrometer. The elemental analysis for nitrogen content was carried out by Kjeldahl method, sulfur content – by Sheniger method [11].

IR-spectra of **IIIa–IIIi**, **IVa–IVg** compounds and  $[\text{Cu}(\text{HL}^4)\text{Cl}_2]_2 \cdot x\text{CH}_3\text{OH}$  ( $x = 0, 2$ ) complex were recorded within  $4000\text{--}400 \text{ cm}^{-1}$  using Specord 75 IR instrument. The samples were prepared as the tablets with KBr.  $^1\text{H}$  NMR spectra were recorded using Varian VXR-200 (200 MHz) and Varian VXR-300 (300 MHz) microwave spectrometers for  $\text{DMSO-d}_6$  solutions with tetramethylsilane as an internal standard.

X-ray analysis of monocrystal **IX** was carried out at 293 K using Siemens P3/PC diffractometer ( $\text{MoK}\alpha$ -radiation, graphite monochromator, scanning method  $q/2q$ ). The whole number of images are 3757, including 3401 independent ones,  $R_{\text{int}} = 0.0740$ . Crystals are monoclinic,  $\text{P}2_1/n$ ;  $a = 8.2001(3) \text{ \AA}$ ,  $b = 20.8052(5) \text{ \AA}$ ,  $c = 11.3421(4) \text{ \AA}$ ;  $\beta = 95.76(3)^\circ$ ,  $V = 1925.31(10) \text{ \AA}^3$ ,  $Z = 2$ . For  $\text{C}_{32}\text{H}_{30}\text{Cl}_4\text{N}_6\text{O}_2\text{S}_2\text{Cu}_2 \cdot 2\text{CH}_3\text{OH}$   $M = 927.74 \text{ g/mol}$ ,  $\rho_{\text{calc.}} = 1.600 \text{ g/cm}^3$ ,  $\mu(\text{MoK}\alpha) = 1.537 \text{ mm}^{-1}$ ,  $F(000) = 948$ .

The components structure was decoded by the direct method and specified by the least-squares method relative to  $F^2$  in full-matrix anisotropic approximation using SHELXTL programs [12]. Hydrogen atoms of methyl groups were assigned geometrically. Two solvated molecules of methyl alcohol are present in the crystal. Full-matrix anisotropic specification of non-hydrogen

atoms is completed at  $R = 0.064$  according to 1918 images with  $I \geq 2\sigma(I)$ ;  $S = 1.103$ .

### 2.2. Synthesis of Hetaryl-2-thiocarboxylic Acid Arylamides

#### 2.2.1. Benzothiazol-2-N-(4-ethoxyphenyl)carbothioamide, IIIc

The reactor equipped by a mechanical stirrer, thermometer and backflow condenser was loaded with 14.9 g (0.1 mol) of 2-methylbenzthiazol, 13.7 g (0.1 mol) of *p*-phenetidine, 9.6 g (0.3 mol) of sulfur and 1.92 g (8 mmol) of  $\text{Na}_2\text{S} \cdot 9\text{H}_2\text{O}$ . The reaction mixture was stirred and 10 ml of benzene was added. Azeotropic solution (benzene+water) was distilled at heating using Dean-Stark head. Then 5 ml of DMSO were added, the temperature was increased to 423 K and the mixture was kept under the mentioned temperature till the end of hydrogen sulfide evolving. The reaction was controlled by the qualitative reaction of lead(II) cation. After the reaction mass was cooled till 333–338 K it was extracted by 5% NaOH solution (3 x 150 ml). The alkaline extracts were collected and neutralized by diluted solution of hydrochloric acid. The precipitate was filtered, dried at the air and recrystallized from the aqueous solutions of methyl and isopropyl alcohols. The yield was 22.6 g (72.0 %); *m.p.* was 400–401 K.

Other arylamides (**IIIa–IIIi**) were synthesized in the same way. Their physico-chemical characteristics are given in Table 1.

#### 2.2.2. Benzimidazol-2-N-(4-bromophenyl)carbothioamide, IVe

The reactor equipped by the mechanical stirrer, thermometer and backflow condenser was loaded with 26.4 g (0.2 mol) of 2-methylbenzimidazole, 36.1 g (0.21 mol) of *p*-bromaniline, 19.2 g (0.6 mol) of sulfur, 3.6 g (15 mmol) of  $\text{Na}_2\text{S} \cdot 9\text{H}_2\text{O}$  and 30 ml of DMSO. The reaction mixture was heated at 383–393 K for 1.5 h without backflow condenser and then at 403–413 K for 10 h with it. The mixture was cooled to 343–353 K and extracted by 5% NaOH (3 x 250 ml). The alkaline extracts were collected. The hot solution was filtrated, cooled to the ambient temperature and acidified by diluted sulfuric acid till pH became 5–6. The yellow precipitate was filtered, dried at the air and recrystallized from the aqueous solution of isopropyl alcohol. Then it was reprecipitated from the diluted NaOH solution and again recrystallized from the aqueous solution of  $\text{CH}_3\text{OH}$ . The yield was 46.8 g (70.5 %); *m.p.* was 432.5–433 K.

Other arylamides (**IVa–IVg**) were synthesized in the same way. Their physico-chemical characteristics are given in Table 2.

### 2.3. Synthesis

#### of $[\text{Cu}(\text{HL}^4)\text{Cl}_2]_2 \cdot x\text{CH}_3\text{OH}$ ( $x = 0, 2$ ) Coordination Compounds

##### 2.3.1. Di( $\mu$ -chloro)-dichloro-bis[benzimidazol-2-N-(4-ethoxyphenyl)carbothioamide]copper(II) solvated by methanol, IX. Method A.

4.46 g (15.0 mmol) of benzimidazol-2-N-(4-ethoxyphenyl) carbothioamide were dissolved in 200 ml of hot unhydrous methyl alcohol. The solution was acidified by 9 ml (90.0 mmol) of 30% HCl and then 0.95 g (15.0 mmol) of copper powder was added under stirring. The obtained mixture was quickly cooled to 293 K and intensively stirred for 11 h. The precipitate of violet-brown color was filtrated using Schott filter, washed by unhydrous methanol ( $3 \times 10$  ml) and dried in desiccator over  $\text{CaCl}_2$ . The yield was 5.96 g (92 %); melting point 473–480 K (with decomp.). Found, %: N 5.98; S 6.57; Cu 13.54. For  $[\text{Cu}_2(\text{C}_{32}\text{H}_{30}\text{N}_6\text{O}_2\text{S}_2)\text{Cl}_4] \cdot 2\text{CH}_3\text{OH}$  calculated, %: N 6.04; S 6.91; Cu 13.70.

##### 2.3.2. Di( $\mu$ -chloro)-dichloro-bis[benzimidazol-2-N-(4-ethoxyphenyl)carbothioamide]copper(II), X. Method B.

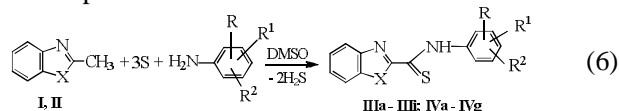
2.59 g (15.2 mmol) of  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$  dissolved in 45 ml of hot unhydrous methyl alcohol and 4.0 ml (40.0 mmol) of 30% HCl were added to 4.46 g (15.0 mmol) of benzimidazol-2-N-(4-ethoxyphenyl) carbothioamide dissolved in 100 ml of hot unhydrous methyl alcohol under constant stirring. The solution was stirred at 318–323 K for 30 min. The formed precipitate of violet-brown color was filtrated using Schott filter, washed by unhydrous methanol ( $3 \times 25$  ml) and dried in desiccator at 363–373 K till the weight became constant. The yield was 6.15 g (95 %); melting point 473–480 K (with decomp.). Found, %: N 9.41; S 7.36; Cu 14.45. For  $[\text{Cu}_2(\text{C}_{32}\text{H}_{30}\text{N}_6\text{O}_2\text{S}_2)\text{Cl}_4] \cdot 2\text{CH}_3\text{OH}$  calculated, %: N 9.73; S 7.43; Cu 14.72.

## 3. Results and Discussion

### 3.1. Synthesis of Heterocyclic Thioamides

To continue the investigations of new heterocyclic thioamides synthesis *via* Willgerodt–Kindler modified reaction [3, 4] we obtained **IIIa–IIIi** and **IVa–IVg**

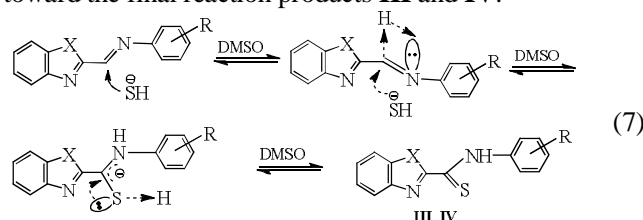
compounds by adding aprotic solvent DMSO to the reaction products:



X = S; R = R<sup>1</sup> = H, R<sup>2</sup> = CH<sub>3</sub>-3, **IIIa**; R = R<sup>1</sup> = H, R<sup>2</sup> = OCH<sub>3</sub>-2, **IIIb**; R = R<sup>1</sup> = H, R<sup>2</sup> = OC<sub>2</sub>H<sub>5</sub>-4, **IIIc**; R = H, R<sup>1</sup> = CH<sub>3</sub>-2, R<sup>2</sup> = CH<sub>3</sub>-4, **III d**; R = H, R<sup>1</sup> = CH<sub>3</sub>-2, R<sup>2</sup> = CH<sub>3</sub>-5, **IIIe**; R = CH<sub>3</sub>-2, R<sup>1</sup> = CH<sub>3</sub>-4, R<sup>2</sup> = CH<sub>3</sub>-6, **III f**; R = R<sup>1</sup> = H, R<sup>2</sup> = Cl-2, **III g**; R = R<sup>1</sup> = H, R<sup>2</sup> = Cl-4, **III h**; R = R<sup>1</sup> = H, R<sup>2</sup> = Br-4, **III i**.  
X = NH; R = H, R<sup>1</sup> = CH<sub>3</sub>-2, **IVa**; R = H, R<sup>1</sup> = OCH<sub>3</sub>-2, **IVb**; R = H, R<sup>1</sup> = OC<sub>2</sub>H<sub>5</sub>-4, **IVc**; R = CH<sub>3</sub>-2, R<sup>1</sup> = CH<sub>3</sub>-4, **IVd**; R = H, R<sup>1</sup> = Cl-4, **IVe**; R = H, R<sup>1</sup> = Br-4, **IVf**; R = H, R<sup>1</sup> = F-4, **IVg**

It was previously noted that  $\text{Na}_2\text{S} \cdot 9\text{H}_2\text{O}$  nucleophilic catalyst activates  $\text{S}_8$  cyclooctasulfone and provides the formation of  $\text{HS}^-$ ,  $\text{HS}_{8-x}\text{S}^-$  reactive nucleophilic ions and final heterocyclic thioamides **III** and **IV** (Eqs. (2)–(5)). The additional introduction of DMSO aprotic solvent increases thioamides yield by 15–27 % due to the following possible reasons:

- the increase in  $\text{HS}_{8-x}\text{S}^-$  ion nucleophilicity;
- the increase in the rate of 2-methylhetarenes methyl group thiolation reaction (formation of intermediate **VI**) and shift of thione-thiol equilibrium toward the final reaction products **III** and **IV**:

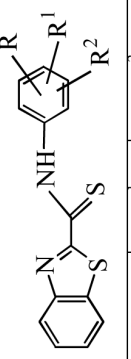


Physico-chemical properties of the synthesized heterocyclic thioamides are represented in Tables 1 and 2. The composition and structure of all synthesized compounds were investigated by elemental analysis and <sup>1</sup>H NMR spectroscopy. In addition, the structure of the compounds **IVa–IVf** was studied by IR-spectroscopy.

So, for the compounds **IVa–IVf** the most typical vibrations are stretching vibrations of  $\nu(\text{N-H})$  thioamide group of average or high intensity in the area of 3370–3260  $\text{cm}^{-1}$ , stretching vibrations of  $\nu(\text{N-H})$  of heterocyclic fragment of average intensity within 3095–3015  $\text{cm}^{-1}$  and multiplex vibrations of thioamide group which were interpreted as vibrations of “B”-, “D”- and “E”-band [7]. The typical stretching vibrations of “B”-band (C=N + N-H) with greater contribution of N-H group; “D”-band (C-N + C=S) with greater contribution of C-N group and “E”-band (C=S + C-N) with greater contribution of C=S group are represented in Table 3.

Table 1

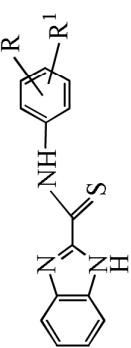
Physico-chemical characteristics of benzothiazoly-2-thiocarboxylic acid arylamides

Compound				Brutto-formula	Found / Calculated, %		<sup>1</sup> H NMR (δ), ppm	Melting point (m.p.), K	Yield*, %
	R	R <sup>1</sup>	R <sup>2</sup>		N	S			
<b>IIIa</b>	H	H	CH <sub>3</sub> -3	C <sub>15</sub> H <sub>12</sub> N <sub>2</sub> S <sub>2</sub>	10.16 / 9.85	22.18 / 22.55	2.40s (3H, CH <sub>3</sub> ); 7.12d, 7.33t, 7.51–7.63m (4H, C <sub>Ar</sub> -H); 7.75t, 8.14d (4H, C <sub>Het</sub> -H); 12.18s (1H, NH)	377–379	17.2
<b>IIIb</b>	H	H	OCH <sub>3</sub> -2	C <sub>15</sub> H <sub>14</sub> N <sub>2</sub> OS <sub>2</sub>	8.94 / 9.33	21.14 / 21.35	3.99s (3H, CH <sub>3</sub> ); 7.03t, 7.17d, 7.31t, 8.65d (4H, C <sub>Ar</sub> -H); 7.54t, 7.61t, 8.13m (4H, C <sub>Het</sub> -H); 11.74s (1H, NH)	379–380	64.8
<b>IIIc</b>	H	H	OC <sub>2</sub> H <sub>5</sub> -4	C <sub>16</sub> H <sub>14</sub> N <sub>2</sub> OS <sub>2</sub>	9.24 / 8.91	20.08 / 20.40	1.38m (3H, CH <sub>3</sub> ); 4.07m (2H, CH <sub>2</sub> ); 6.94d, 7.88d (4H, C <sub>Ar</sub> -H); 7.52t, 7.58t, 8.10d (4H, C <sub>Het</sub> -H); 12.21s (1H, NH)	400–401	72.0
<b>IIIId</b>	H	CH <sub>3</sub> -2	CH <sub>3</sub> -4	C <sub>16</sub> H <sub>14</sub> N <sub>2</sub> S <sub>2</sub>	9.67 / 9.39	21.80 / 21.49	2.22s, 2.35s (6H, 2-CH <sub>3</sub> , 4-CH <sub>3</sub> ); 7.05–7.25m (3H, C <sub>Ar</sub> -H); 7.52–7.62m, 8.10–8.16t (4H, C <sub>Het</sub> -H); 12.12s (1H, NH)	412–414	68.5
<b>IIIe</b>	H	CH <sub>3</sub> -2	CH <sub>3</sub> -5	C <sub>16</sub> H <sub>14</sub> N <sub>2</sub> S <sub>2</sub>	9.05 / 9.39	21.84 / 21.49	2.12s, 2.35s (6H, 2-CH <sub>3</sub> , 5-CH <sub>3</sub> ); 7.08d, 7.17t (3H, C <sub>Ar</sub> -H); 7.54t, 7.60t, 8.10–8.16m (4H, C <sub>Het</sub> -H); 12.10s (1H, NH)	405–406	61.7
<b>IIIf</b>	CH <sub>3</sub> -2	CH <sub>3</sub> -4	CH <sub>3</sub> -6	C <sub>17</sub> H <sub>16</sub> N <sub>2</sub> S <sub>2</sub>	9.24 / 8.97	20.11 / 20.53	2.11s, 2.26s (9H, 2-CH <sub>3</sub> , 4-CH <sub>3</sub> , 6-CH <sub>3</sub> ); 6.90s (2H, C <sub>Ar</sub> -H); 7.55m, 8.06m (4H, C <sub>Het</sub> -H); 12.20s (1H, NH)	448.5–451	80.5
<b>IIIg</b>	H	H	Cl-2	C <sub>14</sub> H <sub>6</sub> ClN <sub>2</sub> S <sub>2</sub>	9.64 / 9.19	20.65 / 21.04	7.38t, 7.43t, 7.59d, 7.94d (4H, C <sub>Ar</sub> -H); 7.54t, 7.60t, 8.09d, 8.15d (4H, C <sub>Het</sub> -H); 12.09s (1H, NH)	441–442	22.0
<b>IIIh</b>	H	H	Cl-4	C <sub>14</sub> H <sub>6</sub> ClN <sub>2</sub> S <sub>2</sub>	9.43 / 9.19	21.75 / 21.04	7.46d, 7.98d (4H, C <sub>Ar</sub> -H); 7.53t, 7.60t, 8.13d, 8.16d (4H, C <sub>Het</sub> -H); 12.40s (1H, NH)	432–434	29.2
<b>IIIi</b>	H	H	Br-4	C <sub>14</sub> H <sub>6</sub> BrN <sub>2</sub> S <sub>2</sub>	8.37 / 8.02	17.96 / 18.36	7.59d, 7.95d (4H, C <sub>Ar</sub> -H); 7.54t, 7.60m, 8.15m (4H, C <sub>Het</sub> -H); 12.37s (1H, NH)	431.5–432.5	51.0

Note: yield of the compounds **IIIa–IIIi** is given after threefold recrystallization according to the following sequence: I – aqueous *i*-C<sub>3</sub>H<sub>7</sub>OH; II – aqueous 5% solution of NaOH; III – aqueous CH<sub>3</sub>OH

Table 2

Physico-chemical characteristics of benzimidazole-2-thiocarboxylic acid arylamides

Compound			Brutto-formula	Found/Calculated, %		<sup>1</sup> H NMR (δ), ppm	Melting point (m.p.), K	Yield*, %
	R	R <sup>1</sup>		N	S			
<b>IVa</b>	H	CH <sub>3</sub> -2	C <sub>15</sub> H <sub>13</sub> N <sub>3</sub> S	15.40 15.72	11.71 11.99	2.30s (3H, CH <sub>3</sub> ); 7.30m, 7.47t, 7.63d, 7.75d (8H, C <sub>Ar</sub> -H, C <sub>Ar</sub> -H); 11.92s, 12.90s (2H, NH, NH)	397–399	54.6
<b>IVb</b>	H	OCH <sub>3</sub> -2	C <sub>15</sub> H <sub>13</sub> N <sub>3</sub> OS	14.60 14.83	11.74 11.32	4.00s (3H, CH <sub>3</sub> ); 7.04t, 7.17d, 7.30t, 7.62d, 7.76d, 8.98d (8H, C <sub>Ar</sub> -H, C <sub>Ar</sub> -H); 11.77s, 12.92s (2H, NH, NH)	444.5–446	60.2
<b>IVc</b>	H	OC <sub>2</sub> H <sub>5</sub> -4	C <sub>16</sub> H <sub>15</sub> N <sub>3</sub> OS	14.51 14.13	10.45 10.78	1.40m (3H, CH <sub>3</sub> ), 4.00–4.10m (2H, CH <sub>2</sub> ); 6.97d, 7.92d (4H, C <sub>Het</sub> -H); 12.05s, 12.85s (2H, NH, NH)	444–445	69.7
<b>IVd</b>	CH <sub>3</sub> -2	CH <sub>3</sub> -4	C <sub>16</sub> H <sub>15</sub> N <sub>3</sub> S	14.54 14.93	11.75 11.40	2.26s, 2.35s (6H, CH <sub>3</sub> , CH <sub>3</sub> ); 7.08t, 7.30t, 7.62d, 7.74d (7H, C <sub>Ar</sub> -H, C <sub>Ar</sub> -H); 11.90s (1H, NH), 12.85s (1H, NH <sub>Het</sub> )	415–417	64.3
<b>IVe</b>	H	Cl-4	C <sub>14</sub> H <sub>10</sub> ClN <sub>3</sub> S	14.88 14.60	11.45 11.14	7.42d, 7.06d, (4H, C <sub>6</sub> H <sub>4</sub> ); 7.28t, 7.68s (4H, C <sub>Het</sub> -H); 12.22s (1H, NH), 12.90s (1H, NH <sub>Het</sub> )	441–442	40.2
<b>IVf</b>	H	Br-4	C <sub>14</sub> H <sub>10</sub> BrN <sub>3</sub> S	12.41 12.65	9.82 9.65	7.59d, 8.00d, (4H, C <sub>Ar</sub> -H); 7.30d, 7.62–7.82m (4H, C <sub>Het</sub> -H); 12.20s, (1H, NH); 12.92s (1H, NH <sub>Het</sub> )	432.5–433.5	70.5
<b>IVg</b>	H	F-4	C <sub>14</sub> H <sub>10</sub> FN <sub>3</sub> S	15.31 15.49	11.42 11.82	2.50d, 3.04s (3H, CH <sub>3</sub> ); 7.27m, 7.68m, (4H, C <sub>Het</sub> -H); 7.55s, 7.88d (3H, C <sub>Ar</sub> -H); 11.80s, (1H, NH); 12.90s (1H, NH <sub>Het</sub> )	446–448	39.9

Note: yield of the compounds **IVa–IVg** is given after threefold recrystallization according to the following sequence: I – aqueous *i*-C<sub>3</sub>H<sub>7</sub>OH; II – aqueous 5% solution of NaOH; III – aqueous CH<sub>3</sub>OH

Table 3

## IR-spectra of thioamides IVa–IVf and complexes IX, X

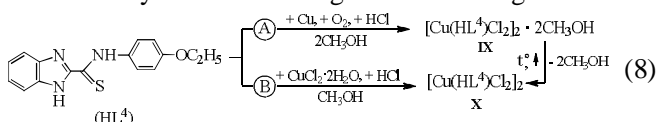
Compound	-NH- group, cm <sup>-1</sup>		-C(=S)NH- group, cm <sup>-1</sup>						Other vibrations, cm <sup>-1</sup>		
	Thioamide group	Benzimidazole fragment	"B"-band			"D"-band				"E"-band	
			C=N	N-H	C-N	C=N	C=S	C=S		C-S	C-N
<b>IVa</b>	3370h	3045a	1555h	1395h	1240a	1085a	850l	755h	1287, 1150, 1240, 950, 775, 880		
	3270h		1510a	1315h	1185a						
<b>IVb</b>	3300h	3063l	1536h	1380h	1280a	1110a	944a	762a	1598, 1465, 1423		
	3260h			1320h				745h			
<b>IVc</b>	3330h	3075a	1532h	1390h	1280a	965a	832a	750h	1225, 1125, 870, 680		
	3260h		1515a	1315a	1185h						
<b>IVd</b>	3324h	3065a	1540h	1387h	1287a	933h	812h	738h	1575, 1146, 900, 700		
	3290h		1506h	1316h	1194h			732h			
<b>IVe</b>	3260h	3015a	1595h	1387h	1290a	1085h	960l	745h	2975, 1455, 772, 870, 835, 695		
			1546h	1320a	1196h						
<b>IVf</b>	3260h	3015l	1596h	1384vh	1280l	1072vh	820h	730vh	1610, 950, 930, 620, 504, 432		
			1535h	1316h		1006a					
<b>IX</b>	3270a	3075a	1545h	1392a	1303a	933a	815a	742h	2935, 2875, 1150, 1435		
	3219a		1520h	1324a	1177h						
<b>X</b>	3270a	3075a	1544h	1392a	1306a	935a	815a	746h	1610, 1435, 1324		
	3210a		1520h		1117h						

Note: "B"-band (C=N + N-H) with greater contribution of N-H group; "D"-band (C-N + C=S) with greater contribution of C-N group and "E"-band (C=S + C-N) with greater contribution of C=S group. The intensity of vibrations: vh – very high; h – high; a – average; l – low

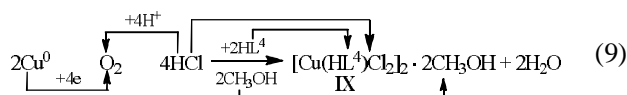
$^1\text{H}$  NMR spectra of the compounds **IIIa–IIIi** (Table 1) represent proton signals of aromatic nuclei, benzothiazole fragment and NH thioamide group. Proton signals of 1,4-disubstituted phenylene (compounds **IIIc**, **IIIh** and **IIIi**) usually appear as two doublets within 6.94–7.98 ppm and signals of 1,2-substituted fragment (**IIIb** and **IIIg**) – as two doublets and two triplets within 7.03–8.65 ppm. Proton signals of methyl group in N-tolylamide fragment (**IIIa**, **IIId**, **IIIe** and **IIIf**) appear as a singlet within 2.11–2.40 ppm. However, protons of the same groups in methoxyl fragment (**IIIb**, **IIIc**) appear as the singlet in a lower field (3.99 ppm), methyl group (**IIIc**) – in the area of 1.38 ppm and methylene group – in the area of 4.07 ppm. Proton signals of benzothiazole fragment usually appear as two doublets (8.06–8.16 ppm) and two triplets (7.52–7.61 ppm). Protons of NH-group appear as the singlet within 12.10–12.40 ppm, except for **IIIb**, the proton of which gives signal at 11.74 ppm. Relation of proton signals of 1,4- and 1,2-disubstituted phenylene, methyl and methoxyl groups of the compounds **IVa–IVg** are similar to those of thioamides **IIIa–IIIi** (Table 2).  $^1\text{H}$  NMR spectra of the compounds **IVa–IVg** differ by two NH-groups in their structure: thioamide and benzimidazole fragments. Signals of NH proton of thioamide group (**IVa–IVg**) appear as the singlet within 11.77–12.22 ppm; protons of benzimidazole fragment – within 12.85–12.92 ppm.

### 3.2. Synthesis of $[\text{Cu}(\text{HL}^4)\text{Cl}_2]_2 \cdot 2\text{CH}_3\text{OH}$ Complex

Earlier we assumed [7] that the coordination compounds obtained *via* the direct synthesis, have a dimeric structure. Their general formula is:  $[\text{Cu}(\text{HL}^{1-3})\text{Cl}_2]_2$ , where  $\text{HL}^1 - \text{C}_7\text{H}_5\text{N}_2\text{C}(=\text{S})\text{NHC}_6\text{H}_5$ ,  $\text{HL}^2 - \text{C}_7\text{H}_5\text{N}_2\text{C}(=\text{S})\text{NHC}_6\text{H}_4\text{CH}_3$ ,  $\text{HL}^3 - \text{C}_7\text{H}_5\text{N}_2\text{C}(=\text{S})\text{NHC}_6\text{H}_4\text{Br}$ . To confirm this assumption here we obtain metal-chelates **IX** and **X** *via* direct and traditional synthesis according to the following scheme:



The same synthesis conditions for the compounds of the general formula  $[\text{Cu}(\text{HL}^{1-3})\text{Cl}_2]_2$  [7] and complex **IX** confirm the identity of their compositions and presence of dimeric structure.  $\text{CH}_3\text{OH}$  solvent and ligands of different nature (thioamide, HCl) may be the potential donors of hydrogen cation. However, neither methyl alcohol (autoprotolysis:  $2\text{CH}_3\text{OH} \rightleftharpoons \text{CH}_3\text{O}^\ominus + \text{CH}_3\text{OH}_2^\oplus$ ), nor thioamide ligand (dissociation of weak NH acid:  $\text{HL} \rightleftharpoons \text{H}^\oplus + \text{L}^\ominus$ ) are capable to compete with the dissociation of strong acid  $\text{HCl} \longrightarrow \text{H}^\oplus + \text{Cl}^\ominus$  and its further participation in formation of compound **IX** dimeric structure:



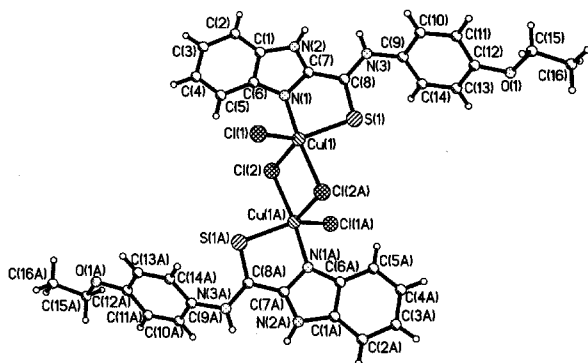
The composition of compounds **IX** and **X** was determined by elemental analysis, their structure – by IR-investigations (Table 3). The comparison of **IVc** and **IX**, **X** IR-spectra confirms the proceeding of complexation reaction and their structure (Scheme (8)).

IR-spectra of the compounds **IX** and **X** are practically identical. There is the shift of thioamide group  $\nu(\text{N-H})$  bond vibrations by  $60\text{ cm}^{-1}$  compared with **IVc** whereas stretching vibrations of benzimidazole fragment  $\nu(\text{N-H})$  bond are the same. For  $\nu(\text{N-H})$  bonds (band “B”) the strong stretching vibrations are typical, as well as their shift toward the high field by 13 and  $9\text{ cm}^{-1}$  for the compounds **IX** and **X**, respectively. Stretching vibrations of  $\nu(\text{C-H})$  and  $\nu(\text{C=S})$  bonds (band “D”) of the average intensity shift toward high and low fields by 23 and  $29\text{ cm}^{-1}$ , respectively. For the band “E” the shift of  $\nu(\text{C=S})$  stretching vibrations toward low field by  $17\text{ cm}^{-1}$  is typical. The spectral data are confirmed by the results of other investigations [7] and prove the structure of the compounds **IX** and **X** obtained *via* both direct and traditional synthesis.

### 3.3. X-ray Analysis of $[\text{Cu}_2(\text{C}_{32}\text{H}_{30}\text{N}_6\text{O}_2\text{S}_2)\text{Cl}_4] \cdot 2\text{CH}_3\text{OH}$ , **IX**

The general view of the molecule, bonds length and valence angles of the compound **IX** are represented in Fig. 1.

Crystal structure of the compound **IX** is a copper(II) bicyclic complex of  $\text{Cu}_2(\text{HL})_2\text{Cl}_4$  ( $\text{HL} = \text{C}_{16}\text{H}_{15}\text{N}_3\text{OS}$ ) with methyl alcohol (1:2). Binuclear complex has two structurally equivalent copper atoms located at the distance of  $3.620\text{ \AA}$  between each other. Every copper atom is five-coordinated by three atoms of chlorine, two of which are bridge ones, and atoms of nitrogen and sulfur of  $(\text{HL}^4)$  thioamide ligand. The coordination polyhedron is a strongly deformed bipyramide, where N(1) and Cl(2A) atoms are in the axial position and S(1), Cl(1) i Cl(2) – in the equatorial one (valence angles N(1)–Cu(1)–Cl(2A)  $170.4(1)^\circ$ ; S(1)–Cu(1)–Cl(1)  $131.54(8)^\circ$ ; S(1)–Cu(1)–Cl(2)  $117.12(7)^\circ$  and Cl(1)–Cu(1)–Cl(2)  $117.12(7)^\circ$ ). The atom of Cu(1) is actually in the plane of equatorial atoms (deviation is  $0.028(1)\text{ \AA}$ ). According to X-ray analysis (Fig. 1) there are three planar fragments in the molecule of thioamide ligand: benzimidazole bicycle (fragment A, mean square deviation of atoms from the plane is  $0.013\text{ \AA}$ ); S(1), N(3), C(8) and C(9) atoms, (fragment B, deviation is  $0.011\text{ \AA}$ ) and ethoxyphenyl radical (fragment C, deviation is  $0.036\text{ \AA}$ ). The rotation angles of fragments B and C relative to A are  $16.8$  and  $111.6^\circ$ , respectively.



Methyl alcohol molecules occupy the position in external sphere and are bounded by hydrogen bonds between each other and nitrogen and oxygen atoms of N(3)–H(3N)···O(1S) thioamide ligand (H···O 1.97 Å, N–H···O 153°).

The length of N(1)–C(7) bond in benzimidazole fragment is 1.341(7) Å, that approximates to the mean value of carbon–nitrogen double bond (1.339 Å) [13]. At the same time C(8)–S(1) bond is shorter than typical values of C=S double bond (cf. 1.632(6) and 1.671 Å). The length of other bonds is typical [14].

## 4. Conclusions

New arylamides of hetaryl-2-thiocarboxylic acid were synthesized *via* Willgerodt–Kindler modified reaction. The introduction of DMSO aprotic solvent into the reaction mass increases the yield of heterocyclic thioamides by 15–27%. The composition and structure of the final products were studied by means of chemical analysis, IR- and <sup>1</sup>H NMR-spectroscopy.

We synthesized di(μ-chloro)-dichloro-bis[benzimidazol-2-N-(4-ethoxyphenyl)carbothioamide]copper(II) coordination compound solvated by methanol. Its composition and binuclear structure were studied using IR-spectroscopy and X-ray analysis.

X-ray analysis was used to examine the dimeric structure of the complex compounds of general formula [Cu(HL<sup>4</sup>)Cl<sub>2</sub>]<sub>2</sub> · xCH<sub>3</sub>OH (x = 0, 2). It was determined that crystals are monoclinic: a = 8.2001(3) Å, b = 20.8052(5) Å, c = 11.3421(4) Å, spatial group P2<sub>1</sub>/*n*, Z = 2. Copper(II) binuclear coordination compound has four chloride anions, two of which are bridge ones, two neutral molecules of benzimidazol-2-N-(4-ethoxyphenyl)carbothioamide (HL), and two solvated molecules of methyl alcohol.

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**Fig. 1.** Molecular structure of the compound IX. The main bonds length (Å): Cu(1)–Cl(1) 2.313(2), Cu(1)–Cl(2) 2.563(2), Cu(1)–Cl(2A) 2.271(2), Cu(1)–S(1) 2.293(2), Cu(1)–N(1) 1.949(4), S(1)–C(8) 1.632(6), O(1)–C(12) 1.362(7), O(1)–C(15) 1.439(7), N(1)–C(7) 1.341(7), N(1)–C(6) 1.396(6). Valence angles: N(1) Cu(1) Cl(2A) 170.4(1)°, N(1) Cu(1) S(1) 84.8(1)°, S(1) Cu(1) Cl(2A) 88.33(7)°, N(1) Cu(1) Cl(1) 94.1(1)°, Cl(1) Cu(1) Cl(2A) 95.55(6)°, S(1) Cu(1) Cl(1) 131.54(8)°, N(1) Cu(1) Cl(2) 89.3(1)°, Cl(2) Cu(1) Cl(2A) 87.78(6)°, S(1) Cu(1) Cl(2) 117.12(7)°, Cl(1) Cu(1) Cl(2) 111.29(6)°, Cu(1) Cl(2) Cu(1A) 92.22(6)°

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## СИНТЕЗ ГЕТЕРОЦИКЛІЧНИХ ТІОАМІДІВ ТА КООРДИНАЦІЙНИХ СПОЛУК КУПРУМУ(II) НА ЇХ ОСНОВІ

**Анотація.** Модифікованою реакцією Вільгеродта-Кіндлера синтезовані ариламиди гетарил-2-тіокарбонової кислоти. Отримані сполуки досліджені методами хімічного аналізу, ІЧ- та <sup>1</sup>H ЯМР-спектроскопією. Методами традиційного та прямого синтезу на основі бензімідазол-2-N-(4-етоксифеніл)карботіоаміда (HL<sup>4</sup>) отримано координаційні сполуки загальної формули [Cu(HL<sup>4</sup>)Cl<sub>2</sub>]<sub>2</sub> · xCH<sub>3</sub>OH (x = 0, 2). Молекулярна і кристалічна структура комплексу [Cu(HL<sup>4</sup>)Cl<sub>2</sub>]<sub>2</sub> · 2CH<sub>3</sub>OH встановлена методом рентгеноструктурного аналізу.

**Ключові слова:** гетероциклічні тіоаміди, біядерні комплекси купруму(II), рентгеноструктурний аналіз.